

Applications of Logarithmic Functions

Logarithmic functions are used to describe scenarios where an exponential change in one quantity corresponds with a linear change in another. Some examples of real-world applications of logarithmic functions include:

- The Richter scale for measuring earthquake intensities.
- The Decibel scale to represent the loudness of sounds.
- The pH scale used in chemistry to measure the acidity of a solution.

Example 1

The amplitude detected from the vibrations of an earthquake increases by a factor of 10 for each increase of 1 on the Richter scale. In other words, an earthquake that measures 7 on the Richter scale has an amplitude that is 10 times larger than one with a measure of 6 on the Richter scale. An earthquake with a reading of 8 on the Richter scale has an amplitude that is 100 times larger than one with a measure of 6 on the Richter scale.

$$\text{Richter Scale Magnitude} = \log(\text{Intensity})$$

← measured in microns (10^{-6} m) of amplitude 100km from the epicenter

How many times more intense is an earthquake with magnitude of 6.8 on the Richter scale compared to one that has a measure of 4.3?

$$\begin{aligned} RS_1 &= \log(\text{Intensity}_1) \\ 6.8 &= \log(\text{Intensity}_1) \\ \text{Intensity}_1 &= 10^{6.8} \end{aligned}$$

$$\begin{aligned} RS_2 &= \log(\text{Intensity}_2) \\ 4.3 &= \log(\text{Intensity}_2) \\ \text{Intensity}_2 &= 10^{4.3} \end{aligned}$$

$$\begin{aligned} \text{Factor} &= \frac{\text{Intensity}_1}{\text{Intensity}_2} \\ &= \frac{10^{6.8}}{10^{4.3}} \\ &= 10^{2.5} \end{aligned}$$

≈ 316 times more intense

Example 2

The loudness of sound is measured in Decibels. For every linear increase of 10 Decibels, the sound intensity increases by a factor of 10. For example, sound that has a loudness of 90 dB is 100 times more intense than sound that has a loudness of 70 dB.

$$\text{Loudness (in dB)} = 10 \log\left(\frac{I}{I_0}\right) \quad \text{where } I_0 = 10^{-12} \text{ W/m}^2$$

How many times more intense is the sound of a rock concert (at 120 dB) compared to normal conversation (at 60 dB)?

$$\begin{aligned} L_{RC} &= 10 \log\left(\frac{I_{RC}}{I_0}\right) \\ 120 &= 10 \log\left(\frac{I_{RC}}{10^{-12}}\right) \\ \frac{120}{10} &= \frac{10 \log\left(\frac{I_{RC}}{10^{-12}}\right)}{10} \\ 12 &= \log\left(\frac{I_{RC}}{10^{-12}}\right) \end{aligned}$$

$$\begin{aligned} \frac{I_{RC}}{10^{-12}} &= \frac{10^{12}}{1} \\ I_{RC} &= 10^0 \text{ W/m}^2 \end{aligned}$$

$$\begin{aligned} L_{NC} &= 10 \log\left(\frac{I_{NC}}{I_0}\right) \\ 60 &= 10 \log\left(\frac{I_{NC}}{10^{-12}}\right) \\ \frac{60}{10} &= \frac{10 \log\left(\frac{I_{NC}}{10^{-12}}\right)}{10} \\ 6 &= \log\left(\frac{I_{NC}}{10^{-12}}\right) \\ \frac{I_{NC}}{10^{-12}} &= \frac{10^6}{1} \end{aligned}$$

$$\begin{aligned} \text{Factor} &= \frac{I_{RC}}{I_{NC}} \\ &= \frac{10^0}{10^{-6}} \\ &= 10^6 \\ \therefore & \text{A million times more intense.} \end{aligned}$$

Example 3

The pH of a solution is a measure of the inverse hydrogen ion $[H^+]$ concentration in units of mol/L. As the pH increases, the amount of hydrogen ions in the solution decreases exponentially. For example, a solution with a pH of 6 is 100 times more acidic than a solution with a pH of 8. Pure distilled water has a pH of 7.

$$\text{pH} = -\log[H^+] \quad \text{where } [H^+] \text{ is the concentration of } H^+$$

a) Determine the pH of hydrochloric acid that is secreted by the stomach lining and has a concentration of 0.03 mol/L.

$$\begin{aligned} \text{pH} &= -\log[H^+] \\ \text{pH} &= -\log(0.03) \\ \text{pH} &= 1.5 \end{aligned}$$

b) Apple juice has a pH of 3. Determine the concentration of hydrogen ions in this solution.

$$\begin{aligned} \text{pH} &= -\log[H^+] \\ 3 &= \frac{-\log[H^+]}{-1} \\ -3 &= \log[H^+] \end{aligned} \quad \begin{aligned} [H^+] &= 10^{-3} \\ [H^+] &= 0.001 \text{ mol/L} \end{aligned}$$

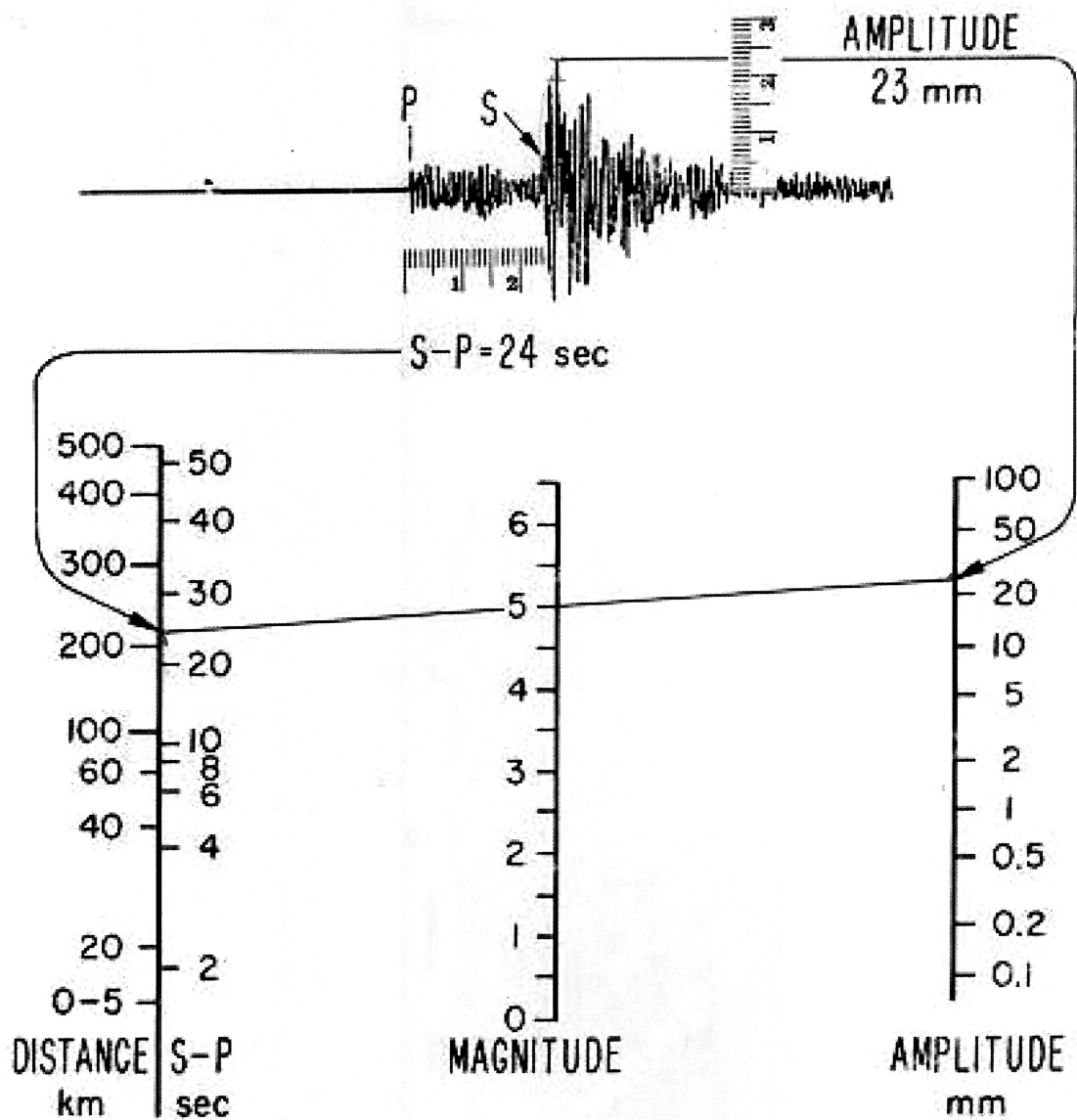
c) If vinegar is 200 times more acidic than tomatoes which has a pH of 4.5, what is the pH of vinegar?

$$\begin{aligned} \textcircled{1} \quad \text{pH}_t &= -\log[H^+]_t \\ 4.5 &= \frac{-\log[H^+]_t}{-1} \\ -4.5 &= \log[H^+]_t \\ [H^+]_t &= 10^{-4.5} \text{ mol/L} \end{aligned}$$

$$\begin{aligned} \textcircled{2} \quad [H^+]_v &= 200 \times [H^+]_t \\ [H^+]_v &= 200 \times 10^{-4.5} \text{ mol/L} \end{aligned}$$

$$\begin{aligned} \textcircled{3} \quad \text{pH}_v &= -\log[H^+]_v \\ \text{pH}_v &= -\log(200 \times 10^{-4.5}) \\ \text{pH}_v &\approx 2.2 \end{aligned}$$

THE RICHTER SCALE



TO DETERMINE THE MAGNITUDE OF AN EARTHQUAKE WE CONNECT ON THE CHART
 A. THE MAXIMUM AMPLITUDE RECORDED BY A STANDARD SEISMOGRAPH, AND
 B. THE DISTANCE OF THAT SEISMOGRAPH FROM THE EPICENTER OF THE
 EARTHQUAKE (OR THE DIFFERENCE IN TIMES OF ARRIVAL OF THE P AND S WAVES)
 BY A STRAIGHT LINE, WHICH CROSSES THE CENTER SCALE AT THE MAGNITUDE